Each question is worth 2 points.

1. Arrange Sr, Cl, and Si in order from lowest ionization energy to highest ionization energy.
   
   _____ < _______ < ________

2. Arrange Se$^{2-}$, Rb$^+$, and As$^{3-}$ in order from the largest ionic radius to the smallest ionic radius.

   _____ > _______ > ________

3. Arrange Rb, P, and Ge in order from the smallest atomic radius to the largest atomic radius.

   _____ < _______ < ________

4. Arrange S, Na, and Al in order from the smallest affinity for electrons to the largest. 
   (The largest affinity for electrons means the one most likely to accept an electron easily.)

   _____ < _______ < ________

5. Arrange Sr, O, and N in order from the one with the smallest electronegativity to the one with the largest electronegativity.

   _____ < _______ < ________
Each question is worth 2 points.

1. Arrange N, K, and Si in order from the smallest atomic radius to the largest atomic radius.
   
   _____ < _______ < ________

2. Arrange S\(^2-\), K\(^+\), and P\(^3-\) in order from the largest ionic radius to the smallest ionic radius.
   
   _____ > _______ > ________

3. Arrange Si, Cl, and Sr in order from largest ionization energy to smallest ionization energy.
   
   _____ > _______ > ________

4. Arrange P, S, and Ba in order from the one with the smallest electronegativity to the one with the largest electronegativity.
   
   _____ < _______ < ________

5. Arrange S, Na, and Al in order from the smallest affinity for electrons to the largest.
   (The largest affinity for electrons means the one most likely to accept an electron easily.)
   
   _____ < _______ < ________
1. Arrange O^{2-}, Na^+, and N^{3-} in order from the largest ionic radius to the smallest ionic radius.

_______ > ________  >  _______

2. Arrange Sr, Cl, and Si in order from lowest ionization energy to highest ionization energy.

_____ < ________ < __________

3. Arrange S, Na, and Al in order from the smallest affinity for electrons to the largest. (The largest affinity for electrons means the one most likely to accept an electron easily.)

_____ < ________ < ________

4. Arrange Se, Cs, and In in order from the largest atomic radius to the smallest atomic radius.

_______ > ________  >  _______

5. Arrange P, S, and Ca in order from the one with the largest electronegativity to the one with the smallest electronegativity.

_____ > ________  >  __________
Name ____________________

CHEM 1364
Thursday Quiz #9
Fall 2011 (Buckley)

Each question is worth 2 points.

1. Arrange P, S, and Ba in order from the one with the largest electronegativity to the one with the smallest electronegativity.
   
   ______ > ________ > _________

2. Arrange S\(^2\)-, K\(^+\), and Cl\(^-\) in order from the largest ionic radius to the smallest ionic radius.
   
   ______ > ________ > _________

3. Arrange O, C, and Ca in order from largest ionization energy to smallest ionization energy.
   
   ______ > ________ > _________

4. Arrange Se, K, and Ga in order from the smallest affinity for electrons to the largest. (The largest affinity for electrons means the one most likely to accept an electron easily.)
   
   ______ < ________ < _________

5. Arrange Rb, P, and Ge in order from the smallest atomic radius to the largest atomic radius.
   
   ______ < ________ < _________